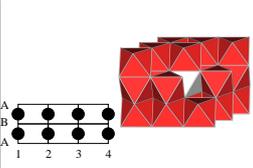
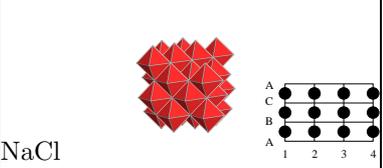
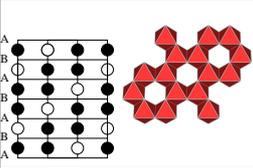
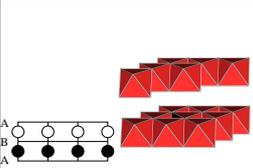
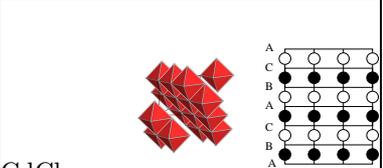
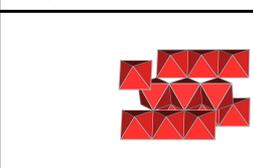
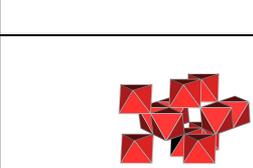
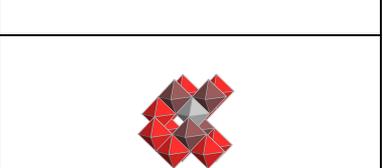
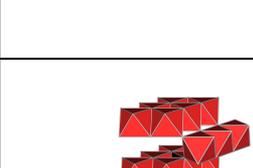
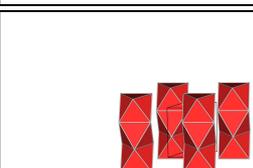


1.1.4. Ionenkristalle (Forts.)

Strukturchemie II: Dichte Packungen mit gefüllten Oktaederlücken

X hexagonal dicht	Formel	CN (X)	Lückenbesetzung	X kubisch dicht
 <p>NiAs</p>	MX	6	1:	 <p>NaCl</p>
 <p>Korund</p>	M <sub>2</sub> X <sub>3</sub>	4	2/3:	-
 <p>CdI<sub>2</sub></p>	MX <sub>2</sub>	3	1/2	 <p>CdCl<sub>2</sub></p>
 <p>CaCl<sub>2</sub> (Rutil)</p>	Raum- netz		1/2:	
 <p>α-PbO<sub>2</sub></p>			1/2:	 <p>Anatas (TiO<sub>2</sub>)</p>
 <p>α-AlOOH</p>			1/2:	
 <p>NFe<sub>2</sub></p>			2/3 u. 1/3	
 <p>ZrI<sub>3</sub></p>	MX <sub>3</sub>	2	1/3	
BiI <sub>3</sub>	Kette	1/3:	2/3 u. 0	YCl <sub>3</sub>
RhF <sub>3</sub>	Schicht			
α-NbI <sub>4</sub> (Kette)	Raumnetz			
α-Nb <sub>2</sub> Cl <sub>10</sub> (mol.)	MX <sub>4</sub>	2/1	1/4	NbF <sub>4</sub> (Schicht)
	MX <sub>5</sub>	2/1	1/5	U <sub>2</sub> Cl <sub>10</sub> (mol.) UF <sub>5</sub> (Kette)
α - WCl <sub>6</sub>	MX <sub>6</sub>	1	1/6	-