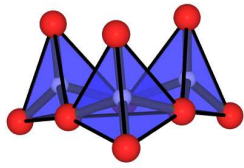
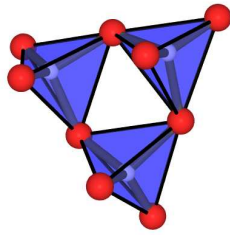


2.6. Chalkogen-Oxide

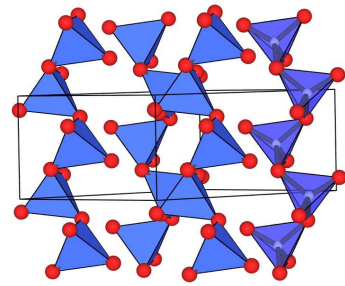
2.6.1. S-Oxide: SO₃



$\gamma - \text{SO}_3$ (eisartig)

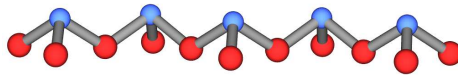


$\gamma - \text{SO}_3$ (eisartig)



$\beta - \text{SO}_3$ (asbestartig)

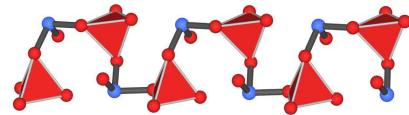
2.6.2. Se-Oxide



SeO₂

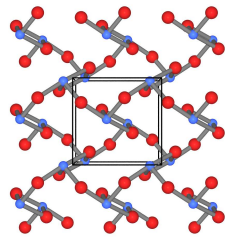


SeO₃

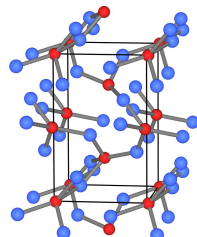


(Se^{VI}O₂O_{2/2})₄ Se₂O₅ (Se^{IV}OO_{2/2})(Se^{VI}O₂O_{2/2})

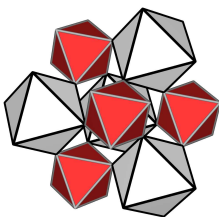
2.6.3. Te-Oxide



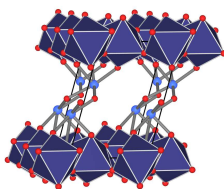
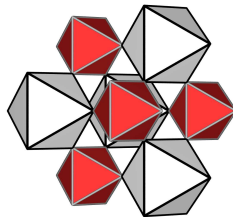
β -TeO₂ (Tellurit)



α -TeO₂



α -TeO₃ (VF₃-Typ + Vergl. mit ReO₃)



Te₂O₅ (Te^{VI}O_{6/2})(Te^{IV}O_{4/2})

2.7. Zusammenfassung kovalente Oxide

	Elemente M	M-Oxid mit maximaler Oxidationsstufe	M-Oxid mit maximaler Oxidationsstufe -2
4. HG 4-bindig 0-LP	C, Si, Ge 	SiO ₂ , GeO ₂ Cristobalit, Tridymit, ...	
5. HG 3-bindig 1-LP	P P, As, Sb 	P ₄ O ₁₀ (H-Form) As ₄ O ₁₀ P ₂ O ₅ (O-Form) As ₂ O ₅	 P ₄ O ₆ As ₂ O ₃ (Claudetit) Sb ₂ O ₃ (Valentinit)
6. HG 2-bindig 2-LP	...	SO ₃ SeO ₃ S S, Se, Te 	
7. HG 1-bindig 3-LP	F ₂ , Cl ₂ , Br ₂ , I ₂ 	Cl ₂ O ₇	 Br ₂ O ₅ I ₂ O ₅
8. HG 0-bindig 4-LP	He, Ne, Ar, Kr, Xe 	XeO ₄	 XeO ₃